

REMARKS

The amendments and remarks herein attend to all outstanding issues in the pending non-final Office Action of October 7, 2003. Claims 1-6, 9, 11-23, 26 and 28-54 remain pending in this application. Claims 1-3, 6, 11-14, 18-20, 28-31, 35-38, 40-43, 45-48 and 50-53 are amended. Claim 7 is cancelled. The Specification is amended for editorial clarity. No new matter is added with any of the amendments presented herein.

I. Priority

The status of related applications has been updated, application serial numbers have been added for the listed issued patents, and wording has been amended to clarify that the present application (and not 08/881,968) is a CIP of 09/419,127. No new matter has been entered.

II. Information Disclosure Statement

Applicants acknowledge that the Examiner is unable to consider references missing publication dates or pages. Regretfully, with the exception of the Merck Manual citation (addressed below), the copies submitted in the Information Disclosure Statement represent the best copies available to Applicant. Applicant is therefore unable to provide the information required for proper consideration.

In accordance with the Examiner's request for page numbers of the Merck Index reference, Applicant submits herewith a Replacement Sheet to the Information Disclosure Statement filed 10/30/2001, and a Replacement Sheet to the Merck manual reference also submitted 10/30/2001. Both Replacement Sheets properly identify pages 1216-1217 of the Merck Manual.

III. Specification

The Merck Index citation was objected to by the Examiner as introducing new matter. The citation has been removed.

IV. Objections to the Claims

All of the Examiner's suggestions regarding changes to the claims have been implemented. The changes were deemed to be semantic in nature and not limiting. No new matter has been entered in the claims.

- The formulas, both structural and written, were corrected in claims 1, 11, 18 and 28.
- The spelling of alkynyl was corrected in claims 1 and 18.
- The "0's have been replaced with "O's in the phosphate salts of claims 1, 11, 18 and 28.
- The term "fungicidal" has been changed to "fungal" in claims 1, 18, 35, 40, 45 and 50.
- The term "salt" has been removed from claims 11, 12, 13, 14 and 18.
- The term "bactericidal" has been changed to "bacterial" in claims 35, 40, 45 and 50.
- The abbreviation AI (active ingredient) was changed to "metal" in claims 2, 3, 19, 20, 38, 43, 48 and 53.
- The lettering of the sub-sections of claims 35, 40, 45 and 50 has been corrected.
- Punctuation has been added at the ends of claims 36, 41, 46 and 51.

In addition to the above amendments, Applicant has further amended the claims for editorial clarity, as follows:

- Tin has been removed from the list of metals in claims 1, 18, 35, 40, 45 and 50.
- The word "has" has been deleted from claim 18.
- Grammar has been corrected in claims 36, 41, 46 and 51.

V. Claim Rejections under 35 U.S.C. §112, first paragraph

Claims 1-6, 9, 11-17, 32 and 33 stand rejected under 35 U.S.C. §112, first paragraph. The Examiner stated that, "the specification, while being enabling for controlling, does not reasonably provide enablement for preventing." Applicant respectfully disagrees with the Examiner's conclusion that prevention is not enabled.

The Merriam-Webster dictionary defines the word prevent in the following manner: "to meet or satisfy in advance; to act ahead of; to keep from happening or existing; to deal with beforehand. Prevent implies taking advance measures against something possible or probable." Therefore, Applicant contends that it would be obvious to one of ordinary skill in the agricultural arts to apply the disclosed composition to an unafflicted plant, thereby "acting ahead of" the fungus, bacteria or microorganism to keep such fungus, bacteria or microorganism from existing on the plant.

The patent application further discloses a minimal formula for a preventative composition, [0024]:

"The preferred composition for preventing Phytophthora will be comprised of at least one potassium phosphonate and at least one potassium phosphate, as it has been found that these two constituents, when combined, will cause a synergistic effect which results in the substantial prevention of infection by Phytophthora...The two constituents will be combined in an amount sufficient to prevent infection and manifestation by various disease causing organisms, with the particular amounts combined dependent upon the particular species of plant to be treated, the specific disease causing organism to be treated, and the particular phosphate salt and phosphonate salt that will be combined."

(emphasis added)

In addition to the formula of a preventative composition (*i.e.*, at least one phosphonate and at least one phosphate), preventative compositions are defined as dry mixes or aqueous solutions [0023] which may be applied by various means: spraying [0032]; foliar application, either by ground or aerial equipment [0038]; or injection or soil applications [0038].

Applicant further submits that there is no requirement that a working example be disclosed in order to enable one of ordinary skill in the art to practice the invention. However, Examples 3, 5, 6, 7 and 8 are all directed to plant material exposed to fungicides or bactericides prior to inoculation or exposure to adverse organisms.

In summary, the application discloses the use of the invention in a preventative sense, a minimal formula for prevention (*i.e.*, at least one phosphonate

and at least one phosphate) and means for administering the preventative composition to plants. It is only left for an ordinary person to know to apply the preventative composition to plants prior to infection. Given that the definition of prevent is "to keep from happening or existing; to deal with beforehand; taking advance measures against something possible or probable," *and* that there are 5 examples of fungicides or bactericides being applied to unafflicted plants in the application, the use of the invention in a preventative sense is clearly enabled.

The preceding discussion demonstrates that the applicant was in intellectual possession of an embodiment using the disclosed composition for preventing fungus, bacteria and microorganisms. Furthermore, no undo experimentation would be necessary for one of ordinary skill in the art to practice this embodiment. For example, sample 7 of Example 13 discloses a composition containing at least one phosphonate and at least one phosphate that would be appropriate for preventative use.

VI. Claim Rejections under 35 U.S.C. §112, second paragraph

Claims 6, 18, 23, 28-31, 35-38, 40, 41-43, 45-48 and 50-53 stand rejected under 35 U.S.C. §112, second paragraph, as indefinite for failing to particularly point out and distinctly claim the subject matter which Applicant regards as the invention. Appropriate correction has been made as follows:

- The recitation of "added as an aqueous solution containing an amount of metal chelate (on a metal basis)" in claims 6 and 23 has been changed to "added in an amount (on a metal basis)". As only one "aqueous solution" is now recited, it is clear that the metal chelate is added on a weight basis relative to the entire solution. Support for this amendment may be found at [0031] which states, "...adding the metal chelate to the water in an amount equal to between about 1% and about 5% by weight (on a metal basis) **of the total solution.**"
- The term "salt" has been removed from claims 11, 12, 13, 14 and 18.

- The terms "first salt" and "second salt" have been changed to "phosphonate salt" and "phosphate", respectively, in claims 28, 29, 30 and 31.
- The word "composition" has been added to claims 35, 38, 40, 43, 48, 50 and 53.
- Potassium phosphate has been changed to "dipotassium" phosphate in claims 36, 37, 41, 42, 46, 47, 51 and 52.
- Phrases directed to the equilibrium reactions and products have been added to claims 35, 40, 45 and 50.

No new matter has been added to the claims by these amendments.

VII. Double Patenting

Claims 1-6, 9, 11-23, 26 and 28-54 stand rejected under the judicially created doctrine of obviousness-type double patenting as being unpatentable over claims 1-8 of U.S. Patent No. 5,736,164, Claims 1-5 of U.S. Patent No. 5,800,837 or Claims 1-49 of U.S. Patent No. 6,338,860, each in view of Ducret et al. (U.S. Patent No. 4,139,616), Fenn et al (1984), Reuveni et al. (Plant Pathology 1995), Scher (U.S. Patent No. 4,714,614) and Supa Crop.

Additionally, Claims 1-6, 9, 11-23, 26 and 28-54 stand rejected under the judicially created doctrine of obviousness-type double patenting as being unpatentable over claims 1 and 2 of U.S. Patent No. 6,139,879 in view of Ducret et al. (U.S. Patent No. 4,139,616), Fenn et al (1984), Reuveni et al. (Plant Pathology 1995), Scher (U.S. Patent No. 4,714,614) and Supa Crop (1990).

Of the above listed references, only Scher and Supa Crop relate to metal chelates. Scher teaches that FeEDDHA is effective against Fusarium wilt. Supa Crop teaches a composition containing phosphates, phosphonates, and metal chelates of iron, manganese, zinc, and copper, wherein the metal chelates provide nutrients to plants. The Examiner has cited *In re Kerkhoven*, 205 USPQ 1069, 1072 (CCPA 1980) which states, “It is prima facie obvious to combine two compositions each of which is **taught by the prior art to be useful for the same purpose**, in order to form a third composition to be used for the very same purpose.” However, Supa Crop

teaches that metal chelates act as nutritional supplements when mixed with a phosphate and a phosphite, whereas Scher teaches fungicidal activity of FeEDDHA combined with a disease-suppressant strain of *Pseudomonas putida*. The two references are not teaching the use of metal chelates for the same purpose. Therefore, the combination of these references does not create a case of *prima facie* obviousness. As none of the other references cited by the Examiner teach metal chelates, they also do not support a case for *prima facie* obviousness-type double patenting.

In addition, as discussed in the next section, there is a synergistic effect based on the combination of Applicant's three-component mixture. This was previously presented to the Examiner. With all due respect, Applicant submits that the Examiner has misinterpreted the supplied information.

VIII. Claim Rejections under 35 U.S.C. §103(a)

Claims 1-6, 9, 11-23, 26 and 28-54 stand rejected under 35 U.S.C. §103(a) as unpatentable over Horriere et al. in view of Ducret et al. (U.S. Patent No. 4,139,616), Fenn et al. (1984), Reuveni et al. (Plant Pathology 1995), Scher (U.S. Patent No. 4,714,614) and Supa Crop (1990).

Applicants repeats and incorporates the arguments set forth in the preceding section. Without more, the pending application claims are in condition for allowance. The fact that the three-component mixture of the invention displays a synergistic effect leaves no room for doubt on this score.

A synergistic effect based on the combination of Applicant's three component mixture was previously presented. In response to these arguments, the Examiner has stated, "In an IPER (12/13/2002), it was remarked by the PCT examiner that the embodiments of the claimed invention showed no synergistic activity but only what would be expected from the combination of the components."

The PCT examiner cited Colby, S. R. "Calculation of the Synergistic and Antagonistic Responses of Herbicide Combinations," *Weeds*, 1967, 15, 20-22, containing the formula:

E=X+Y-XY/100

in which E represents the expected percentage of inhibition of the disease for the combination of the two fungicides at defined doses (for example equal to x and y respectively), X is the percentage of inhibition observed for the disease by the compound (I) at a defined dose (equal to x), Y is the percentage of inhibition observed for the disease by the compound (II) at a defined dose (equal to y). **When the percentage of inhibition observed for the combination is greater than E, there is a synergistic effect.** The PCT Examiner then incorrectly states, "this is the case neither in Exhibit A, nor in Exhibit B, wherein the percentage of inhibition in both cases equals to E (more or less 1-2%)." A comparison of expected results (calculated with the above formula) versus observed results for Exhibits A and B are shown below:

Exhibit A – Late Blight Observations on Potato var. "Atlantic"

Treatment	% Late Blight
Sample 1. 15.1% K ₂ HPO ₃ + 15.1% K ₂ HPO ₄ + 69.8% inert ingredients (H ₂ O)	1.85
Sample 2. Cu-EDDHA (3.2% Cu) – 0.2 lb. AI/100 gal.	18.45
Sample 7. 80% v/v Sample 1 + 0.2 lb. AI/100 gal. (Cu-EDDHA)	0.39
Control	28.12

E = (28.12-1.85)+(28.12-18.45)-((28.12-1.85)*(28.12-18.45))/100 = 33.4% inhibition expected (i.e., 33.4-28.12 = **5.28% late blight expected**)

0.39% late blight observed

Improvement of 4.89%

Exhibit B – Carrot Alternaria Trial

Treatment	% Infected
Sample 1. 15.1% K ₂ HPO ₃ + 15.1% K ₂ HPO ₄ + 69.8% inert ingredients (H ₂ O)	26.25
Sample 2. Cu-EDDHA (3.2% Cu) – 0.2 lb. AI/100 gal.	28.75
Sample 7. 80% v/v Sample 1 + 0.2 lb. AI/100 gal. (Cu-EDDHA)	15.0
Control	52.5

E = (52.5-26.25)+(52.5-28.75)-((52.5-26.25)*(52.5-28.75))/100 = 36.14% inhibition
 expected (i.e., 52.5 – 36.14 = **16.36% infection expected**)

15.0% infection observed

Improvement of 1.36%

Additionally, data based on an affidavit submitted in the present application on July 3, 2003 is shown below:

Cauliflower seedlings subjected to Zanthomonas campestris pv. campestris

Treatment	% Infected
FNX-100. (PO ₃ /PO ₄) @ 2% v/v (H ₂ O)	21.0
FNX-200. Cu-EDDHA @ 0.2 lb. AI/100 gal.	34.7
FNX-2000. FNX-100@2% + FNX-200 0.2 lb. AI/100 gal. (Cu-EDDHA)	1.39
Control	59.3

E = (59.3-21.0)+(59.3-34.7)-((59.3-21.0)*(59.3-34.7))/100 = 53.48% inhibition
 expected (i.e., 59.3 – 53.48 = **5.82% infection expected**)

1.39% infection observed

Improvement of 4.43%

Each of the three examples above shows greater than expected inhibition of disease and, thus, a synergistic effect. The court has stated, “A greater than expected result is an evidentiary factor pertinent to the legal conclusion of obviousness...of the claims at issue.” In *re Corkill*, 711 F.2d 1496, 226 USPQ 1005 (Fed. Cir. 1985).

Improvements in pesticide effectiveness, such as those demonstrated above, are not insignificant. World pesticide expenditures totaled more than \$33.5 billion dollars in 1998 and 1999, with U.S. expenditures making up approximately \$11 billion dollars of total expenditures.¹ Therefore, improvements in pesticide efficiency of 1-5% translate into savings of billions of dollars.

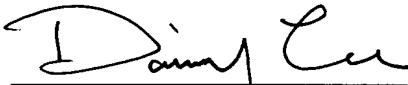
¹ "Pesticides Industry Sales and Usage: 1998 and 1999 Market Estimates" issued by the United States Environmental Protection Agency in August 2002, accessed January 27, 2004 on the internet at: www.epa.gov/oppbead1/pestsales/99pestsales/market_estimates1999.pdf.

VIII. Conclusion

Applicant submits that the present paper is a complete response to the non-final Office Action of October 7, 2003 that should put this application in condition for allowance. Should the Examiner have any questions, comments or suggestions that would expedite the prosecution of this case, Applicants' undersigned representative earnestly requests a telephone conference.

Respectfully Submitted,

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use: In analytical chemistry.

THERAP CAT: Vasodilator; antidiote (cyanide poisoning).
7636. Potassium Nitroprussiate. Dipotassium pentakis(cyano-Nitroso)ferrate(2-); potassium nitrophenylpentacyanoferrate(II). $C_5FeK_2N_6O_7$; mol wt 294.14. C 20.42%, Fe 18.95%, K 26.58%, O 54.44%. $K_2Fe(CN)_5NO$. Dihydrate, garnet-red, hygroscopic crystals. Sol in 1 part water, in alcohol. Keep well closed.

7637. Potassium Oleate. Oleic acid potassium salt. Approx $C_{18}H_{34}KO$. Yellowish or brownish, soft mass. Sol in water, alc. The aq soln is alkaline to phenolphthalein.

USE: Detergent.

7638. Potassium Osmate(VI). $K_2O_8O_3$; mol wt 332.40. K 23.53%, O 19.25%; Os 57.22%. K_2OsO_4 .

Dihydrate, violet, hygroscopic crystals. Poisonous! Sol in water, insol in alcohol, ether. Slowly dec in aq solns with formation of the tetroxide. Keep well closed.

7639. Potassium Oxalate. $C_2K_2O_4$; mol wt 166.22. C 14.45%, K 47.05%, O 38.50%. $K_2C_2O_4$.

Monohydrate, colorless, odorless crystals; efflorescent in warm dry air. Poisonous! d 2.13. Loses its water at about 160°, when ignited is converted into carbonate without appreciable charring. Sol in 3 parts water.

USE: Cleaning and bleaching straw, removing stains in photography; in examination of blood to prevent its coagulation; also in analytical chemistry.

7640. Potassium Percarbonate. $C_4K_2O_6$; mol wt 198.22. C 12.12%, K 39.45%; O 48.43%. K_2CO_3 . Prepn of practically anhydrous compd: Partington, Fathallah, J. Chem. Soc. 1950, 1934.

Monohydrate, white, granular mass. Sol in water with evolution of oxygen. One part potassium percarbonate is sol in 15 parts of cold water; dec in boiling water; 100 parts water dissolve 6.5 parts potassium percarbonate at ordinary temp. Keep dry and protected from light.

USE: Has been used in microscopy for detecting tubercle bacilli stained with ruthenium in smears in photography under the name "Anti-hypo," to remove last traces of sodium thiosulfate; also as oxidizing agent in chem analyses, but is no longer favored. Caution: Strong irritant. Causes vomiting if swallowed. Large quantities can be fatal.

7641. Potassium Perchlorac. Potassium Perchlorate. $KClO_4$; mol wt 138.55. Cl 25.59%, K 28.22%; O 46.19%.

Colorless crystals or white, crystall powder. Dec at 400°; also due by organic matter, oxidizable substances and on concussion, but is less reactive than the chlorate. d 2.52. Sol in 65 parts cold water, 15 parts boiling water; practically insol in alcohol.

USE: In explosives, pyrotechnics and photography, in analytical chemistry.

7642. Potassium Periodate. Potassium metaperiodate. IKO_4 ; mol wt 230.01. I 55.18%; K 17.00%; O 27.82%.

Prepd by oxidizing potassium iodate with chlorine in alkaline soln. Hill, J. Am. Chem. Soc. 50, 2678 (1928); Inorg. Syn. 1, 171 (1939).

Colorless tetragonal crystals, d¹⁵ 3.618. mp 382°. Sol in water (G/100 g H₂O): 0.168 at 0°; 0.42 at 20°; 0.93 at 40°; 2.16 at 60°; 4.44 at 80°; 7.87 at 100°; also given as 0.66 at 13°.

Sparingly sol in aq KOH.

USE: Powerful oxidizer in acid soln, oxidizing manganese compounds to permanganate; used for this purpose in analytical chemistry (colorimetric estimation of Mn), also for the oxidation of some organic compds. Caution: Highly irritating to skin, eyes, mucous membranes.

7643. Potassium Permanganate. Permanganic acid potassium salt; channelon mineral. $KMnO_4$; mol wt 158.03. K 24.74%; Mn 34.76%; O 40.50%. Prepn from manganese one by electrolytic oxidation: Faith, Keyes & Clark's Industrial Chemicals, F. A. Lowenthal, M. K. Moran, Eds. (Wiley Interscience, New York, 4th ed., 1975) pp. 679-683.

Dark purple or bronze-like, odoreless crystals. Almost opaque by transmitted light and of a blue metallic luster by reflected light. Sweet with astrigent aftertaste; stable in air.

use: In buffering agent. d 1.1. soluble in water, 3.5 parts cold, 3.5 parts boiling water. Dec by alc with many other organic solvents, also by concd acids with loss of oxygen; with HCl, chlorine is liberated. Readily decomposed by reducing substances, such as ferrous salts, iodates, etc., especially in the presence of an acid. Caution: Take great care in handling as explosions may occur if brought into contact with organic or other readily oxidizable substances, either in soln or in the dry state. LD ₅₀ orally rats: 1.09 g/kg; H. F. Smyth et al., Am. Ind. Hyg. Assoc. 30, 470 (1969).	7649. Potassium Phosphate, Tribasic. Tripotassium phosphate. K_3PO_4 ; mol wt 212.27. K 55.15%; O 30.15%. P 14.8%. K_2PO_4 ; mol wt 193.55%; K ₂ HPO ₄ ; mol wt 145.98%. Simon, Schulze, Z. Anorg. Allg. Chem. 127, 75 (1927). Simon, Schulze, Z. Anorg. Allg. Chem. 142, 331 (1939).	7650. Potassium Phosphate, Tetrabasic. Tetraphosphoric acid, orthorhombic crystals, d ¹⁵ 2.364. mp 134°. Soln in water: 43.7% at 0°; 50.8% at 25°; 59.7% at 65.1°. Insol in alcohol. Aq solns are strongly alkaline. Octahydrate, flat, rectangular platelets, mp 43°.
use: In buffering agent. d 1.1. soluble in water, 3.5 parts cold, 3.5 parts boiling water. Dec by alc with many other organic solvents, also by concd acids with loss of oxygen; with HCl, chlorine is liberated. Readily decomposed by reducing substances, such as ferrous salts, iodates, etc., especially in the presence of an acid. Caution: Take great care in handling as explosions may occur if brought into contact with organic or other readily oxidizable substances, either in soln or in the dry state. LD ₅₀ orally rats: 1.09 g/kg; H. F. Smyth et al., Am. Ind. Hyg. Assoc. 30, 470 (1969).	7651. Potassium Picrate, 2,4,6-Trinitrophenol potassium picrate(V). Approx $K_3H_2Sb_2O_7 \cdot 4H_2O$. White granules or cryst powder. Moderately sol in cold water; dec by heat. Very sol in water; insol in alcohol. Keep well closed.	7652. Potassium Pyrophosphate, Acid. Potassium antimonate(V). Approx $K_3H_2Sb_2O_7 \cdot 4H_2O$. White granules or cryst powder. Formerly used for detection of sodium.
use: Topical antiseptic, astringent, dorant.	7653. Potassium Pyrophosphate. Diposphoric acid tetrasodium salt; potassium trinitrophenolate. $C_6H_4K_2N_3O_4$; mol wt 270. C 15.73%; N 15.75%; O 41.92%. K 16.3%. Potassium Sorbate, 2,4-Hexaum salt; sorbic acid potassium salt; potas BB powder. $C_6H_4K_2O_4$; mol wt 1:1. Tetrahydrate, translucent crystals or cooling saline taste. Slightly effloresces mp 70-80°; at 100° loses H_2O , becomes impure. Acid, calcium or lead salts begin to dec. Sol in 0.9 part w alcohol. The aq soln is slightly alkaline silver nitrate.	7654. Potassium Persulfate. Disulfuric acid dipotassium salt; "anhydrous" potassium acid sulfate. $K_2O_8S_2$; mol wt 330.34. K 47.35%. O 47.35%. $K_2P_2O_7$; mol wt 308.5. O 18.75%. $K_2P_2O_7$; Manuf: Faith, Keyes & Clark's Industrial Chemicals, F. A. Lowenthal, M. K. Moran, Eds. (Wiley Interscience, New York, 4th ed., 1975) pp. 684-687. Sol in alcohol. The aq soln is acidic. Keep well closed, in a cool place. Use: Bleaching fabrics, soaps; in photography under safelight; colorants; deliquescent granules or crystal mass; in detergents and surfactants; in water treatment; in thinning mud as a clay thinner.
use: Topical antiseptic, astringent, dorant.	7655. Potassium Pyrosulfate. Disulfuric acid dipotassium salt; potassium acid sulfate. $K_2O_8S_2$; mol wt 330.34. O 47.35%. $K_2S_2O_8$; mol wt 328.2. Sol in water; insol in alcohol. The aq soln is strongly acid.	7656. Potassium Pyrosulfate. Disulfuric acid dipotassium salt; potassium acid sulfate. $K_2O_8S_2$; mol wt 330.34. O 47.35%. $K_2S_2O_8$; mol wt 328.2. Sol in water; insol in alcohol. The aq soln is strongly acid.
use: Topical antiseptic, astringent, dorant.	7657. Potassium Salicylate. Salicylic acid potassium salt; potassium phenacetin. $C_7H_6KO_3$; mol wt 176.21. Sol in water; insol in alcohol. The aq soln is practically neutral to litmus. Sol in water, alcohol. The aq soln is strongly alkaline.	7658. Potassium Salicylate. Salicylic acid potassium salt; potassium phenacetin. $C_7H_6KO_3$; mol wt 176.21. Sol in water; insol in alcohol. The aq soln is strongly alkaline.
use: Topical antiseptic, astringent, dorant.	7659. Potassium Silver Cyanide. no- <i>Chargentate(I-)</i> ; potassium dicyanogen. C_2AgCN_2 ; mol wt 190.21%. Ag 54.21%; N 14.65%; K 14.08%. White crystals; sensitive to light. Pois acids. In silver plating; as bactericide.	7660. Potassium Sodium Tartrate. genetic salt. $C_4H_4KNaO_6$; mol wt 211.1.2%. K 18.60%; Na 10.94%; O 45.68%. Tetrahydrate, translucent crystals or cooling saline taste. Slightly effloresces mp 70-80°; at 100° loses H_2O , becomes impure. Acid, calcium or lead salts begin to dec. Sol in 0.9 part w alcohol. The aq soln is slightly alkaline silver nitrate.
use: Topical antiseptic, astringent, dorant.	7661. Potassium Sorbate, 2,4-Hexaum salt; sorbic acid potassium salt; potas BB powder. $C_6H_4K_2O_4$; mol wt 1:1. Tetrahydrate, colorless crystals. d ¹⁵ 1.363. Crystals. d ¹⁵ 1.363. Dec above 270°.	7662. Potassium Stannate(IV). K_2SnO_3 ; mol wt 319.15%. O 19.60%; Sn 48.47%. Trihydrate, colorless crystals. d 3.1. Water, where greater solv in water is desirable where greater solv in water is desirable.
use: As mold and yeast inhibitor.	7663. Potassium Stannosulfate. $M_2SnS_4O_6$; mol wt 389.02. K 20.10%; O 30.51%; $K_2Sn(SO_4)_2$. White crystals. Partly dec by water; insol in alcohol. The aq soln is:	7664. Potassium Stearate. Stearic Acid. Approx $KC_{18}H_{35}O_2$. The article of c considerable proportion of palmitate. White powder; usually with slight od in cold, readily in hot water or alc used in textile dying and printing.
use: Topical antiseptic, astringent, dorant.	7665. Potassium Sulfate. Salicylic acid potassium salt; potassium phenacetin. $C_7H_6KO_3$; mol wt 176.21. Sol in water; insol in alcohol. The aq soln is:	7666. Potassium Sulfate. Salicylic acid potassium salt; potassium phenacetin. $C_7H_6KO_3$; mol wt 176.21. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7667. Potassium Sulfite. K_2SO_3 ; mol wt 174.15. K 22.19%; H 2.86%. Sol in water; insol in alcohol. The aq soln is:	7667. Potassium Sulfite. K_2SO_3 ; mol wt 174.15. K 22.19%; H 2.86%. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7668. Potassium Selenite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:	7668. Potassium Selenite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7669. Potassium Selendite. K_2Se ; mol wt 174.18. Sol in water; insol in alcohol. The aq soln is:	7669. Potassium Selendite. K_2Se ; mol wt 174.18. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7670. Potassium Selenite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:	7670. Potassium Selenite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7671. Potassium Selenite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:	7671. Potassium Selenite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7672. Potassium Phosphate, Dibasic. Dipotassium hydrogen phosphate. DKP; dipotassium hydrogen phosphate. DKP ; K_2HPO_4 ; mol wt 174.18. H 0.57%. Sol in water; insol in alcohol. The aq soln is:	7672. Potassium Phosphate, Dibasic. Dipotassium hydrogen phosphate. DKP; dipotassium hydrogen phosphate. DKP ; K_2HPO_4 ; mol wt 174.18. H 0.57%. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7673. Potassium Phosphate, Monobasic. K_2HPO_4 ; mol wt 174.18. H 0.57%. Sol in water; insol in alcohol. The aq soln is:	7673. Potassium Phosphate, Monobasic. K_2HPO_4 ; mol wt 174.18. H 0.57%. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7674. Potassium Phosphate, Trisodium salt; potassium trisodium phosphate. K_3PO_4 ; mol wt 212.27. Sol in water; insol in alcohol. The aq soln is:	7674. Potassium Phosphate, Trisodium salt; potassium trisodium phosphate. K_3PO_4 ; mol wt 212.27. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7675. Potassium Salicylate. Salicylic acid potassium salt; potassium phenacetin. $C_7H_6KO_3$; mol wt 176.21. Sol in water; insol in alcohol. The aq soln is:	7675. Potassium Salicylate. Salicylic acid potassium salt; potassium phenacetin. $C_7H_6KO_3$; mol wt 176.21. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7676. Potassium Salicylate. Salicylic acid potassium salt; potassium phenacetin. $C_7H_6KO_3$; mol wt 176.21. Sol in water; insol in alcohol. The aq soln is:	7676. Potassium Salicylate. Salicylic acid potassium salt; potassium phenacetin. $C_7H_6KO_3$; mol wt 176.21. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7677. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:	7677. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:
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use: Topical antiseptic, astringent, dorant.	7679. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:	7679. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7680. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:	7680. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7681. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:	7681. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7682. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:	7682. Potassium Selendite. K_2SeO_4 ; mol wt 221.15. K 22.19%; O 44.87%; S 36.13%; K_2SO_4 ; mol wt 140.02. Sol in water; insol in alcohol. The aq soln is:
use: Topical antiseptic, astringent, dorant.	7683. Potassium Silicate. Soluble potassium silicate. K_2SiO_4 ; mol wt 140.02. Sol in water. Insol in alcohol. pH 7.	7683. Potassium Silicate. Soluble potassium silicate. K_2SiO_4 ; mol wt 140.02. Sol in water. Insol in alcohol. pH 7.

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